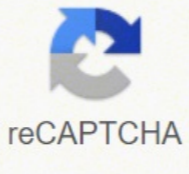
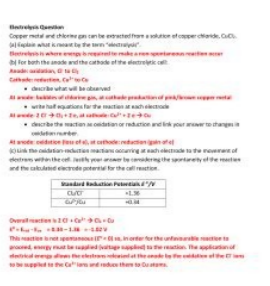




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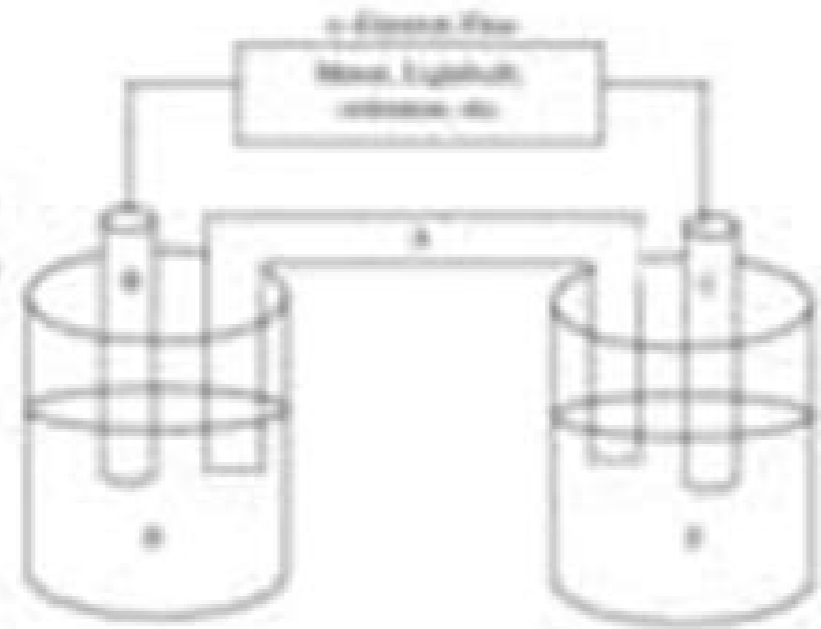


Honors Chemistry Electrochemistry Problems

Name _____

Date _____ Period _____

The cell will show a voltage as long as the metals Zn(s) and Cu(s) are in contact with their ions. Answer the following questions. Pay attention to the direction of electron flow. Answer the following questions using the labels indicated above. The solutions may be used over and over.



1. Salt bridge
2. Voltmeter
3. Cathode (Zn electrode)
4. Cathode (Cu electrode)
5. Anode (Zn electrode)
6. Anode
7. Cathode
8. Electrode where oxidation occurs
9. Electrode where reduction occurs
10. Zn(s)
11. Electrode that gets smaller with time
12. Electrode that gets larger with time
13. Write reduction potentials for Zn and Cu ions. $E^\circ_{Zn} = -0.76V$ and $E^\circ_{Cu} = +0.34V$ respectively, what is the voltage of the cell?

Write a balanced cell diagram for this reaction.

Write the oxidation half-reaction for the cell.

Write the reduction half-reaction for the cell.

For the following four questions identify the half-reaction as either oxidation or reduction.

14. $Fe^{2+} \rightarrow Fe^{3+} + e^-$
15. $Cl_2 \rightarrow 2Cl^- + 2e^-$
16. $Cl_2 + 2e^- \rightarrow 2Cl^-$
17. $2H_2 \rightarrow 2H_2^+ + 2e^-$
18. $2H^+ + 2e^- \rightarrow H_2$
19. For the reaction $2Fe^{3+} + 3I^- \rightarrow 2Fe^{2+} + 3I_2$ which species is the reducing agent? Explain.

Factsheet live Transport

Fig. 108 - Diagram of a mammalian nephron.

PDF

Observations:

- 1. The glomerulus is a cluster of capillaries.
- 2. The Bowman's capsule is a double-walled structure that surrounds the glomerulus.
- 3. The proximal convoluted tubule is the first part of the renal tubule.
- 4. The loop of Henle is a U-shaped part of the renal tubule.
- 5. The distal convoluted tubule is the second part of the renal tubule.
- 6. The collecting duct is a tube that carries urine away from the nephron.

Name _____ Date _____ Block _____

Lab: Redox Half-Reactions

Purpose: To determine the oxidation and reduction half-reactions for simple single-replacement reactions.

Observations:
Record your observations below (e.g., gas bubbles or color change). If nothing occurs, write "No Rxn."

	1	2	3	4	5	6
	ZnSO ₄	H ₂ SO ₄	FeSO ₄	AgNO ₃	CuSO ₄	MgSO ₄
A Zn						
B Cu						
C Mg						
D Fe						

Analysis:
For each reaction that took place:
i. Write a complete and balanced chemical equation, including phases.
ii. Assign oxidation numbers for every element.
iii. Write an the oxidation and reduction half-reactions, including phases.

Example:

Overall Rxn:	$Ba(s) + H_2SO_4(aq) \rightarrow H_2(g) + BaSO_4(aq)$
Oxidation numbers:	0 +1 +6 -2 0 +2 +6 -2
Oxidation half Rxn:	$Ba(s) \rightarrow Ba^{2+}(aq) + 2e^-$
Reduction half Rxn:	$2H^+(aq) + 2e^- \rightarrow H_2(g)$

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