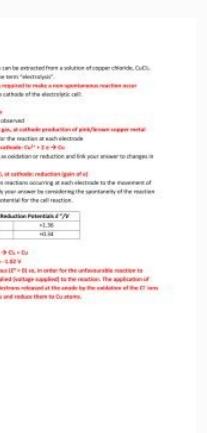


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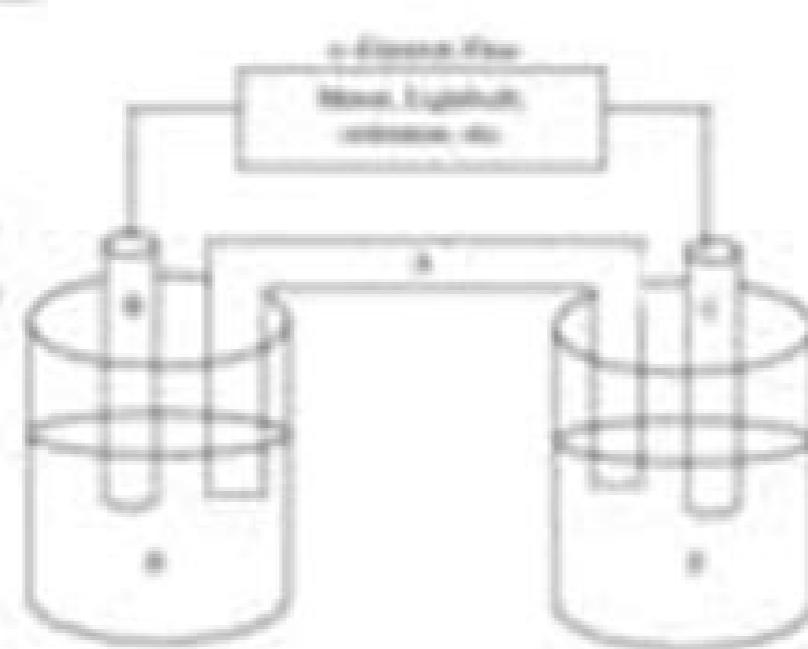
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Honors Chemistry Electrochemistry Problems

Name _____ Date _____ Period _____

The cells will allow direct comparison of the standard reduction potentials of Zinc and Copper. It is important that the two half-reactions have the same oxidation state. The electrons in the electron flow must be balanced.



1. Salt bridge
2. Zinc electrode
3. Zinc ions (Zn²⁺)
4. Zinc ions (Zn)
5. Zinc electrode
6. Anode
7. Cathode
8. Zinc ions (Zn²⁺) solution
9. Zinc ions (Zn²⁺) solution
10. CuSO₄
11. Zinc/Zinc ions (Zn²⁺) half-reaction
12. Zinc/Zinc ions (Zn²⁺) half-reaction

(1) Write reduction potentials for Ni and Cu as -0.27V and +0.34V respectively. what is the voltage of the cell?

Write a balanced half-reaction for this reaction.

Write the oxidation half-reaction for the cell.

Write the reduction half-reaction for the cell.

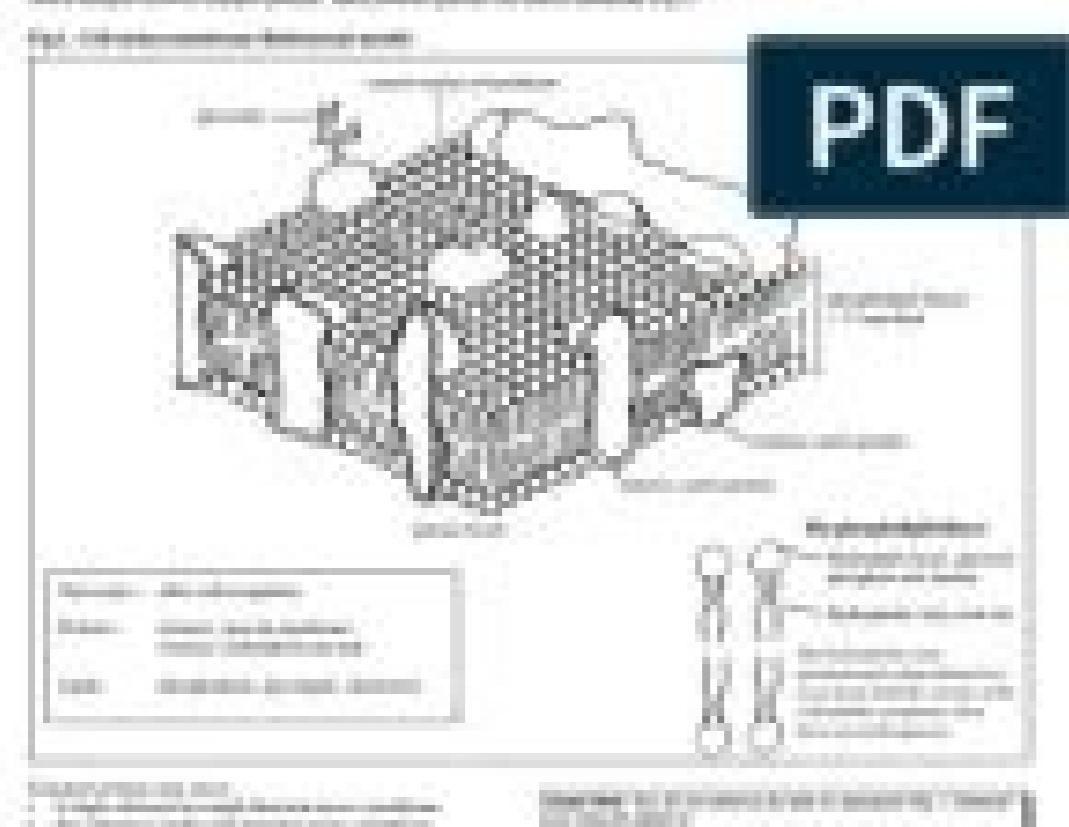
For the following four questions identify the half-reaction as either oxidation or reduction:

10. $\text{MnO}_4^- \rightarrow \text{Mn}^{2+} \rightarrow \text{Mn}$
 11. $\text{Cl}_2 \rightarrow \text{Cl}^{2-} \rightarrow \text{Cl}$
 12. $\text{Cr}_2\text{O}_7^{2-} \rightarrow \text{Cr}^{3+} \rightarrow \text{Cr}$
 13. $\text{Br}_2 \rightarrow \text{Br}^{2-} \rightarrow \text{Br}$
 14. $\text{Zn}^{2+} \rightarrow \text{Zn} \rightarrow \text{Zn}$
15. Redox reaction:
 $\text{2MnO}_4^- + \text{3Zn} \rightarrow \text{2Mn}^{2+} + \text{4Zn}^{2+}$
 which species is the reducing agent? Explain.

Factsheet 4

live Transport

PDF



Name _____ Date _____ Block _____

Lab: Redox Half-Reactions

Purpose: To determine the oxidation and reduction half-reactions for simple single-replacement reactions.

Observations

Record your observations below (e.g., gas bubbles or turn black). If nothing occurs, write "No Rxn".

	1	2	3	4	5	6
ZnSO ₄	ZnSO ₄	H ₂ SO ₄	FeSO ₄	AgNO ₃	CuSO ₄	MgSO ₄
A Zn	1	2	3	4	5	6
B Cu	7	8	9	10	11	12
C Mg	13	14	15	16	17	18
D Fe	19	20	21	22	23	24

Analysis

For each reaction that took place:

- i. Write a complete and balanced chemical equation, including phases.
- ii. Assign oxidation numbers for every element.
- iii. Write out the oxidation and reduction half-reactions, including phases.

Overall Rxn:	$\text{Ba(s)} + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{H}_2\text{S}(g) + \text{BaSO}_4(\text{aq})$
Oxidation numbers:	0 +1 +6 -2 0 +2 +6 -2
Oxidation half Rxn:	$\text{Ba(s)} \rightarrow \text{Ba}^{2+}(\text{aq}) + 2\text{e}^-$
Reduction half Rxn:	$2\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{H}_2\text{S}(g)$

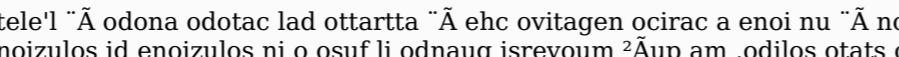
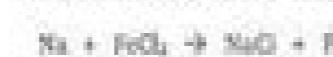
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Redox Reactions:

ASSIGNING OXIDATION NUMBERS

Assign oxidation numbers to all of the elements in each of the compounds or ions below.

1. HCl	11. H_2SO_4
2. KNO ₃	12. H_2SO_3
3. OH ⁻	13. BaO ₂
4. MgN ₃	14. KMnO ₄
5. KOH	15. NH
6. AgNO ₃	16. MnO ₂
7. I ₂	17. OF ₂
8. H ₂ O ₂	18. SO ₃
9. PbO ₂	19. NH ₃
10. NaHSO ₄	20. Na



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